

ContainerPower Energy Solutions

Vanadium flow battery structure



Overview

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ed network. Flow batteries (FB) store chemical energy and generate electricity by a redox reaction between vanadium ions dissolved in the electrolytes. FB are essentially comprised of two key elements (Fig. 1): the cell stacks, where chemical energy is converted to electricity in a reversible.

A vanadium flow battery works by circulating two liquid electrolytes, the anolyte and catholyte, containing vanadium ions. During the charging process, an ion exchange happens across a membrane. This process changes the oxidation states of the vanadium ions, leading to efficient electricity.

Redox flow batteries store the energy in the liquid electrolytes, pumped through the cell and stored in external tanks, rather than in the porous electrodes as for conventional batteries. This approach offers interesting solutions for low-cost energy storage, load leveling and power peak shaving.

In VFBS, this electrolyte is composed of vanadium dissolved in a stable, non-flammable, water-based solution. Vanadium is a non-toxic, widely-available metal that is typically used for making steel more ductile, strengthening titanium and even as a dietary supplement. Vanadium is produced globally.

The vanadium redox flow battery (VRFB) is one promising candidate in large-scale stationary energy storage system, which stores electric energy by changing the oxidation numbers of anolyte and catholyte through redox reaction. This chapter covers the basic principles of vanadium redox flow.

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